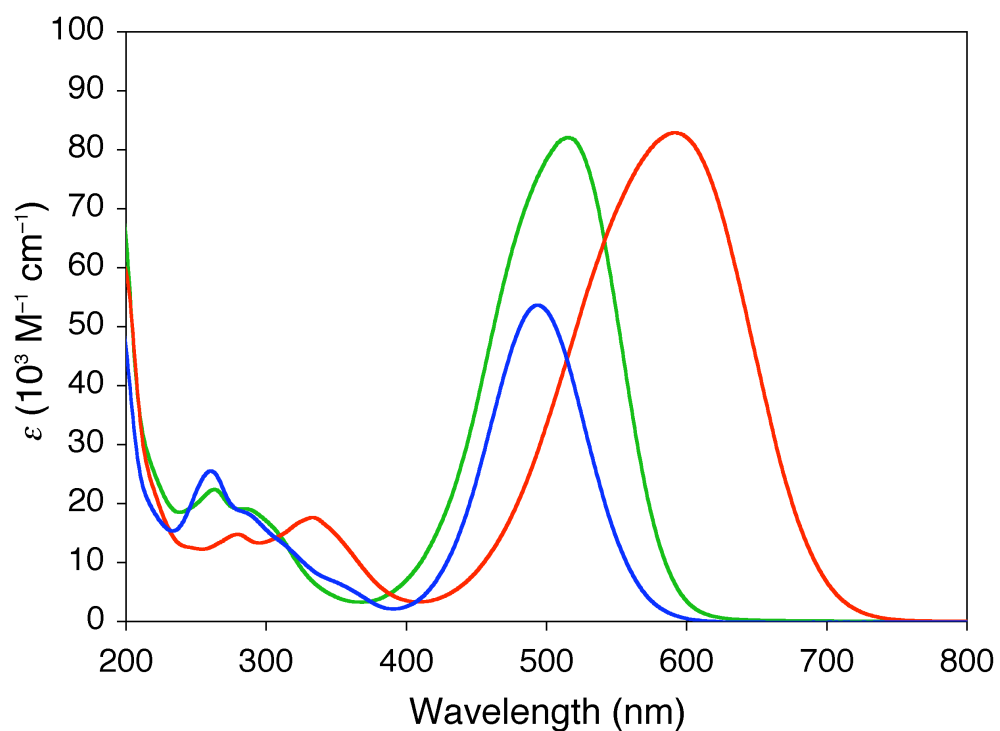


# Syntheses and Properties of Two-Dimensional, Dicationic Nonlinear Optical Chromophores Based on Pyrazinyl Cores

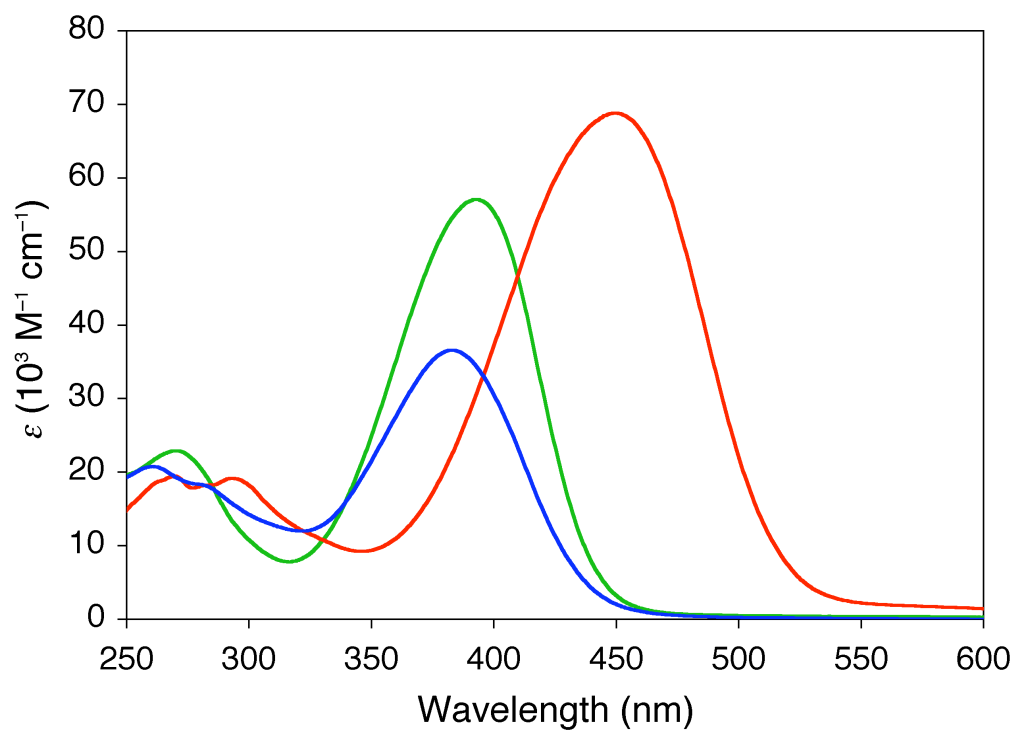
Benjamin J. Coe, John Fielden, Simon P. Foxon, Madeleine Helliwell, Inge Asselberghs, Koen Clays, Kurt De Mey and Bruce S. Brunschwig

## Supporting Information

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**Figure S1.** Electronic absorption spectra of the  $-NMe_2$ -substituted series  $[1][PF_6]_2$  (blue),  $[2][PF_6]_2$  (green) and  $[3][PF_6]_2$  (red) in acetonitrile at 293 K.

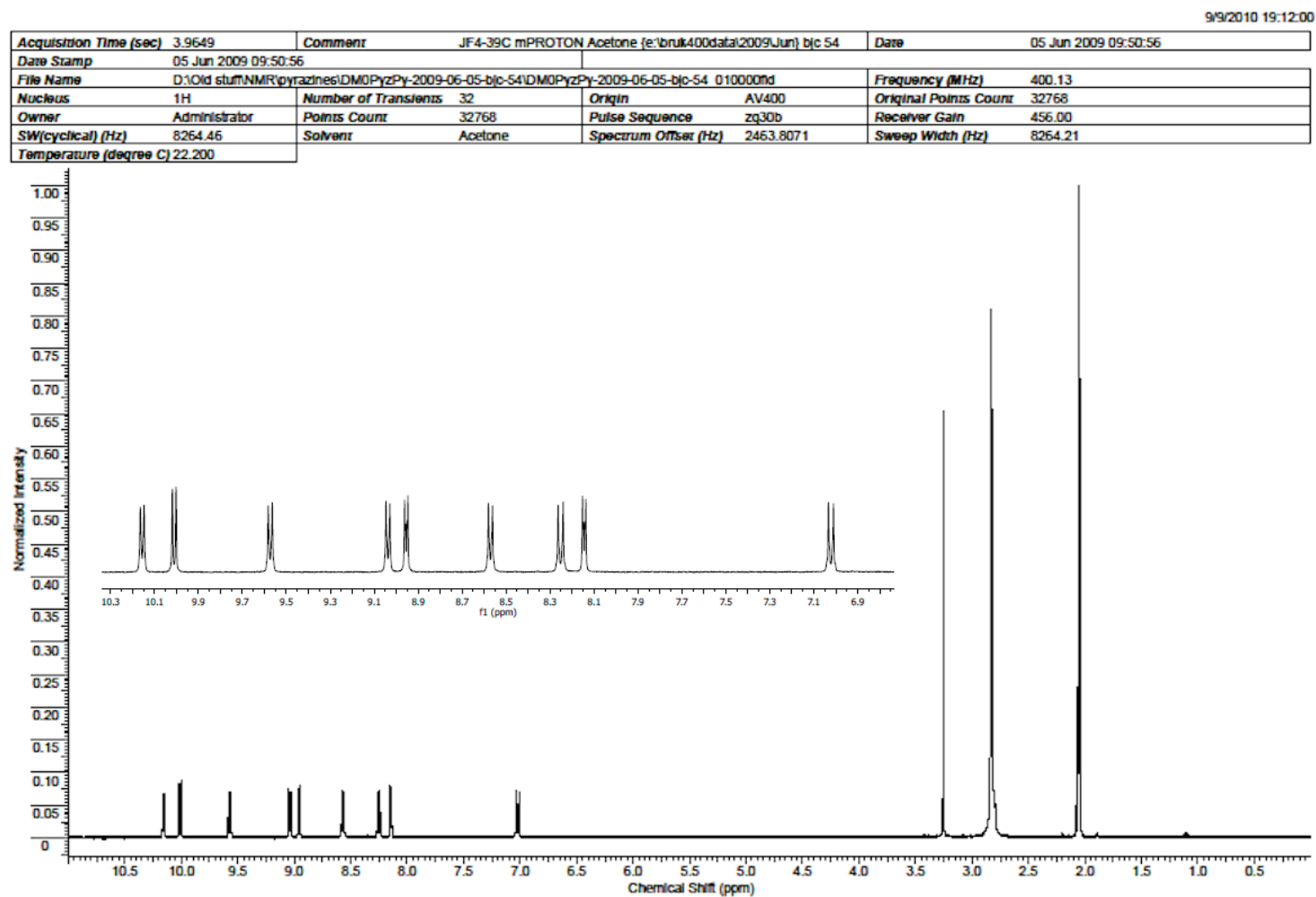


**Figure S2.** Electronic absorption spectra of the  $-OMe$ -substituted series  $[4][PF_6]_2$  (blue),  $[5][PF_6]_2$  (green) and  $[6][PF_6]_2$  (red) in acetonitrile at 293 K.

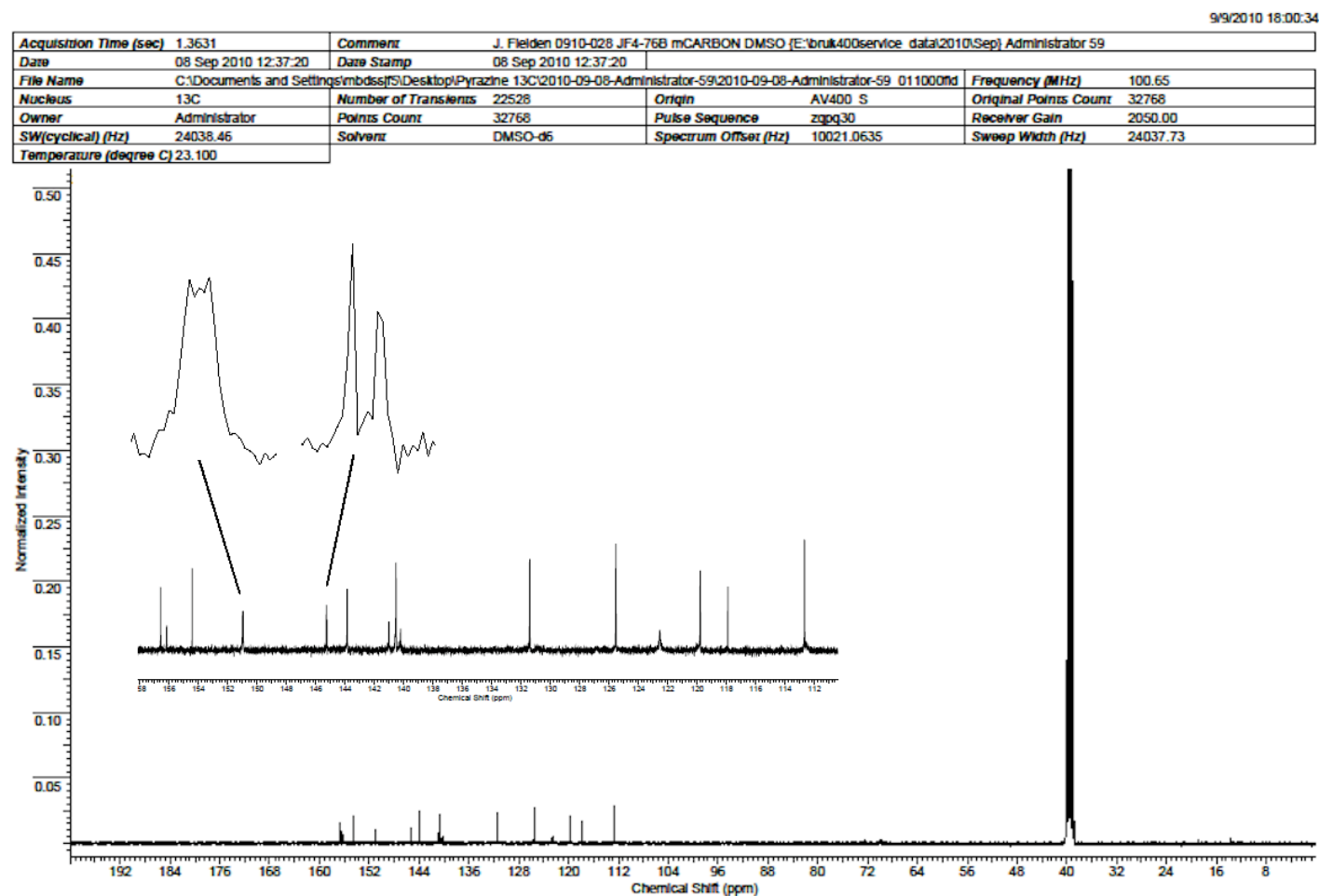
## Original NMR Spectra

Over the following pages NMR spectra ( $^1\text{H}$  in  $(\text{CD}_3)_2\text{CO}$  and  $^{13}\text{C}$  in  $(\text{CD}_3)_2\text{SO}$ ) are presented for all of the new hexafluorophosphate salts described in this paper. Only  $^1\text{H}$ -NMR spectra are shown for the three nitrate salts (in  $\text{CD}_3\text{OD}$ ), which were isolated in smaller quantities primarily for the purpose of obtaining X-ray crystal structures. It should be noted that the relative solubility of the samples is generally moderate to low, and the molecular weights are high, so many of the  $^{13}\text{C}$  spectra and some of the  $^1\text{H}$  spectra show only rather weak signals. All spectra are shown with the full range recorded (0–11 ppm for  $^1\text{H}$  and 0–200 ppm for  $^{13}\text{C}$ ), with expansions to show more clearly the aromatic environments. Note that the methyl signals from the  $-\text{NMe}_2$ -based chromophores are obscured by the DMSO solvent signals.

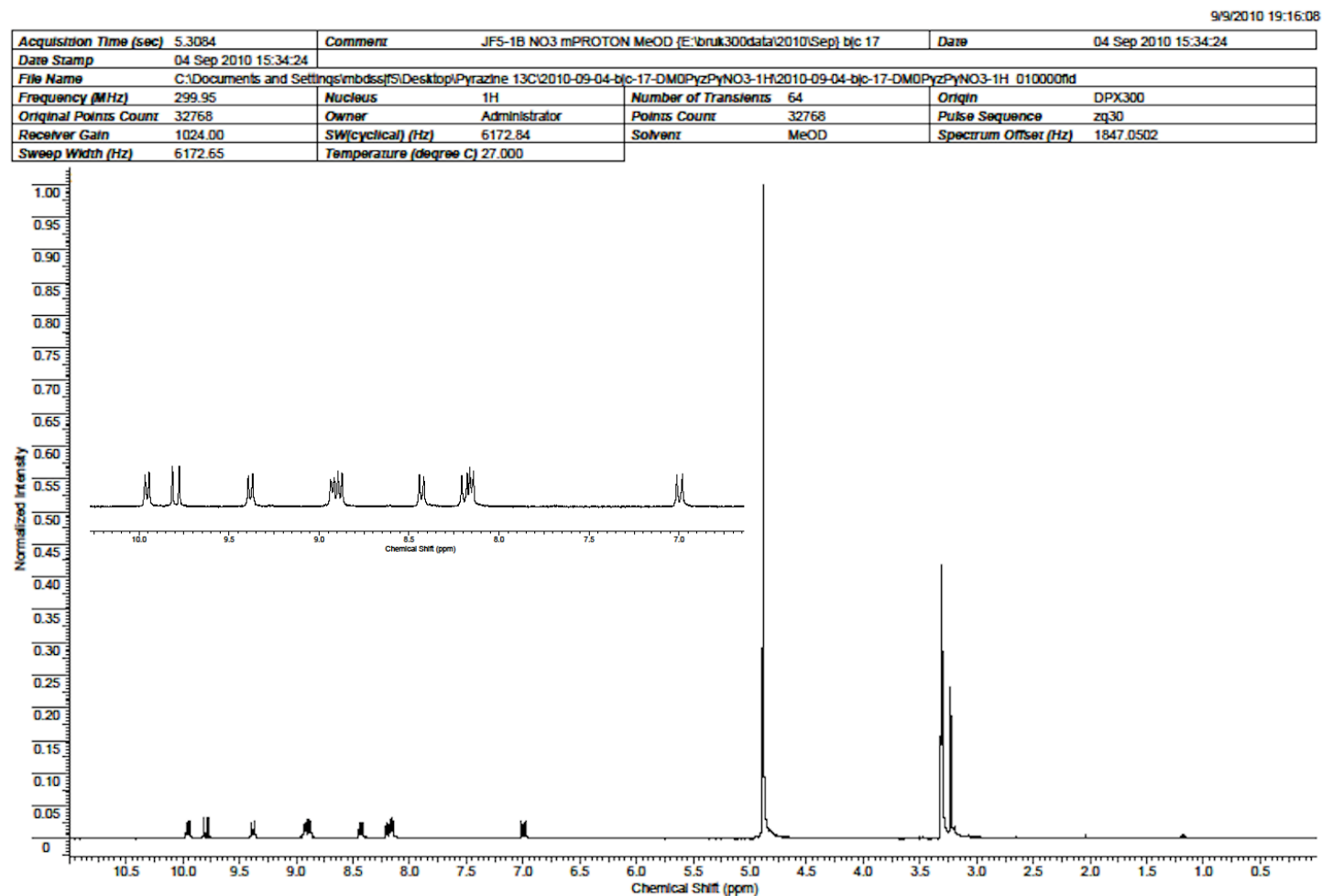
**Figure S3.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{1}][\text{PF}_6]_2$  (400 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).



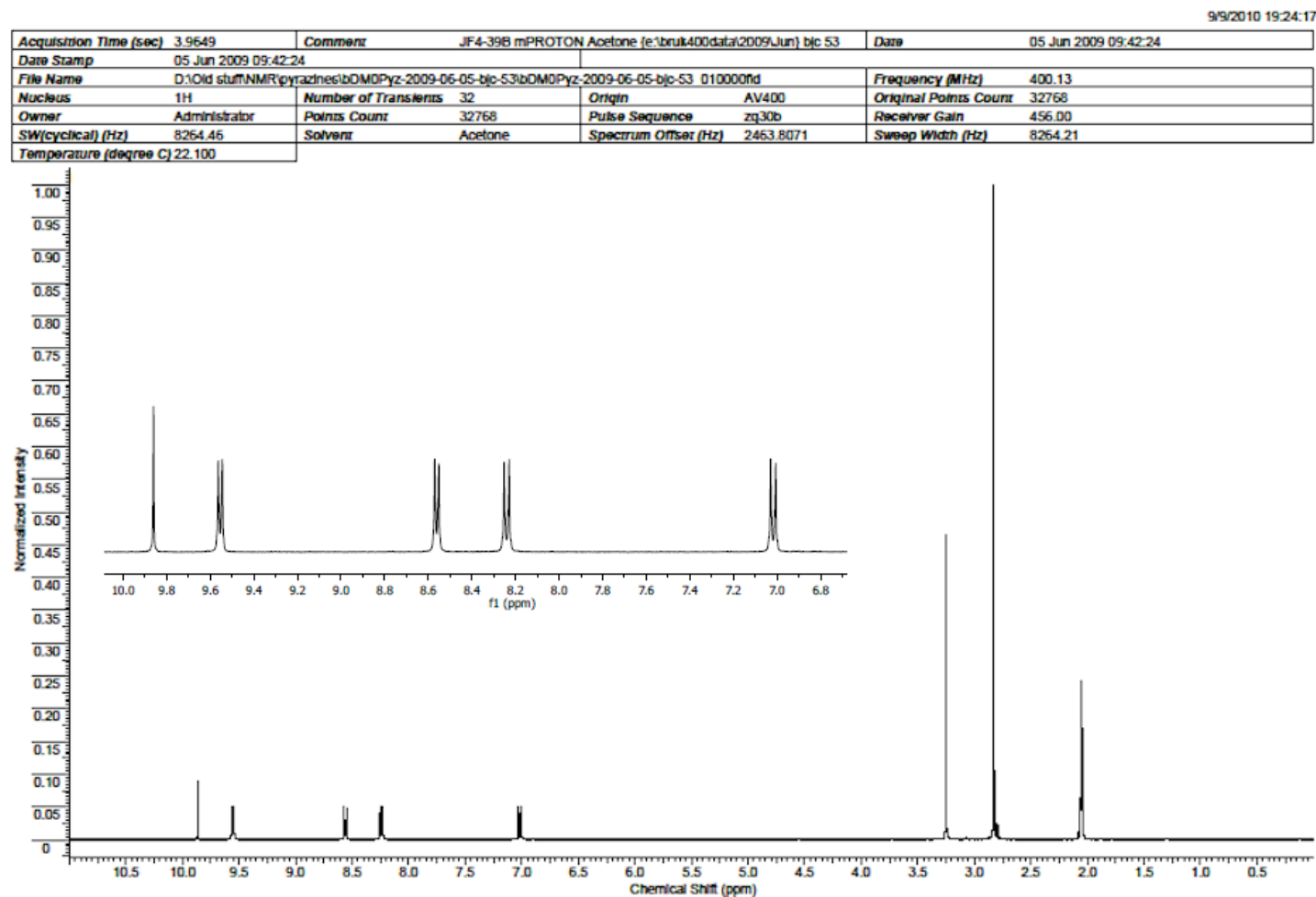
**Figure S4.**  $^{13}\text{C}$  NMR spectrum of  $[\mathbf{1}][\text{PF}_6]_2$  (100 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K)  
(extra expansions show the closely spaced environments at 150.99/150.96 and 145.28/145.25 ppm)



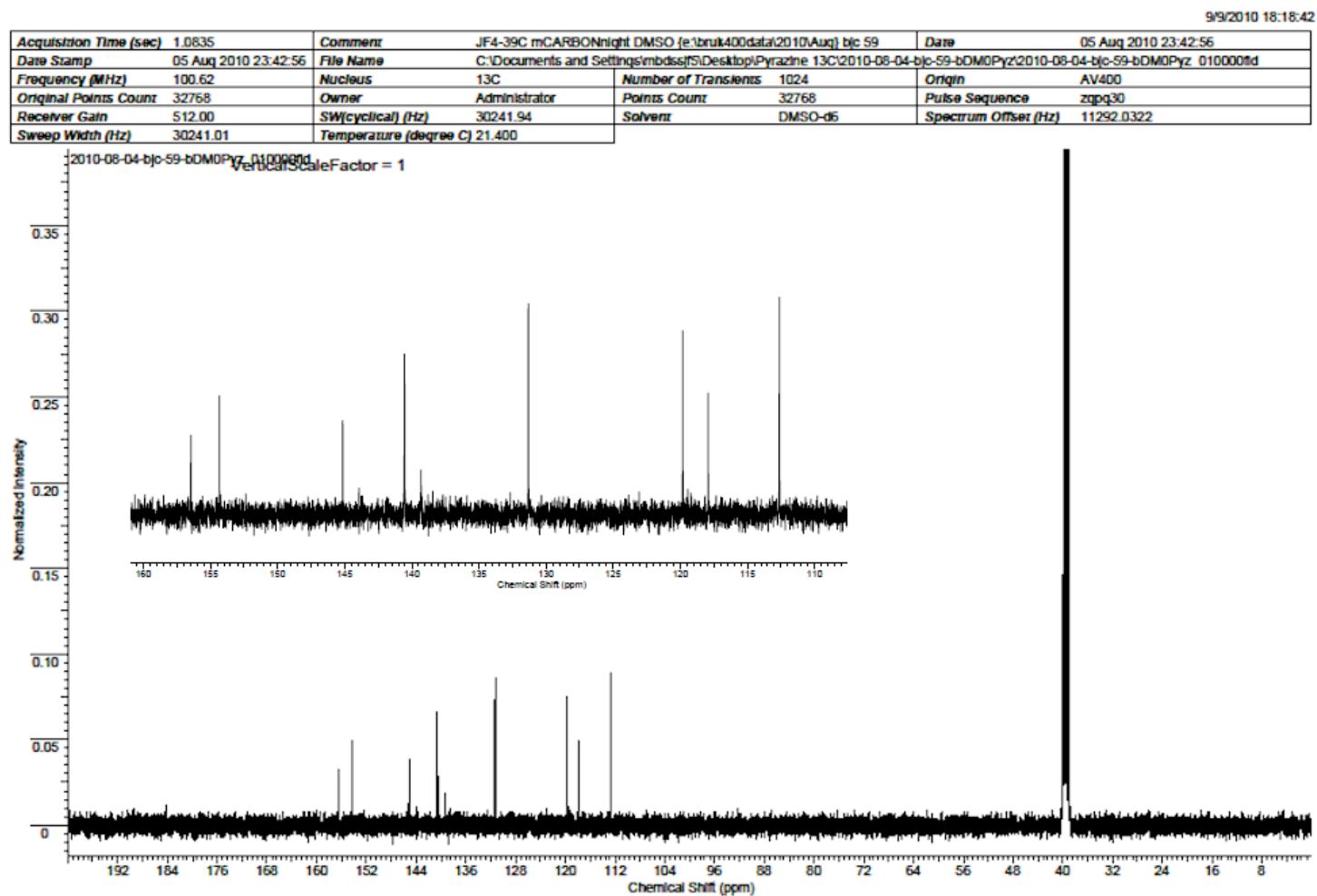
**Figure S5.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{1}][\text{NO}_3]_2$  (300 MHz in  $\text{CD}_3\text{OD}$  at 293 K).



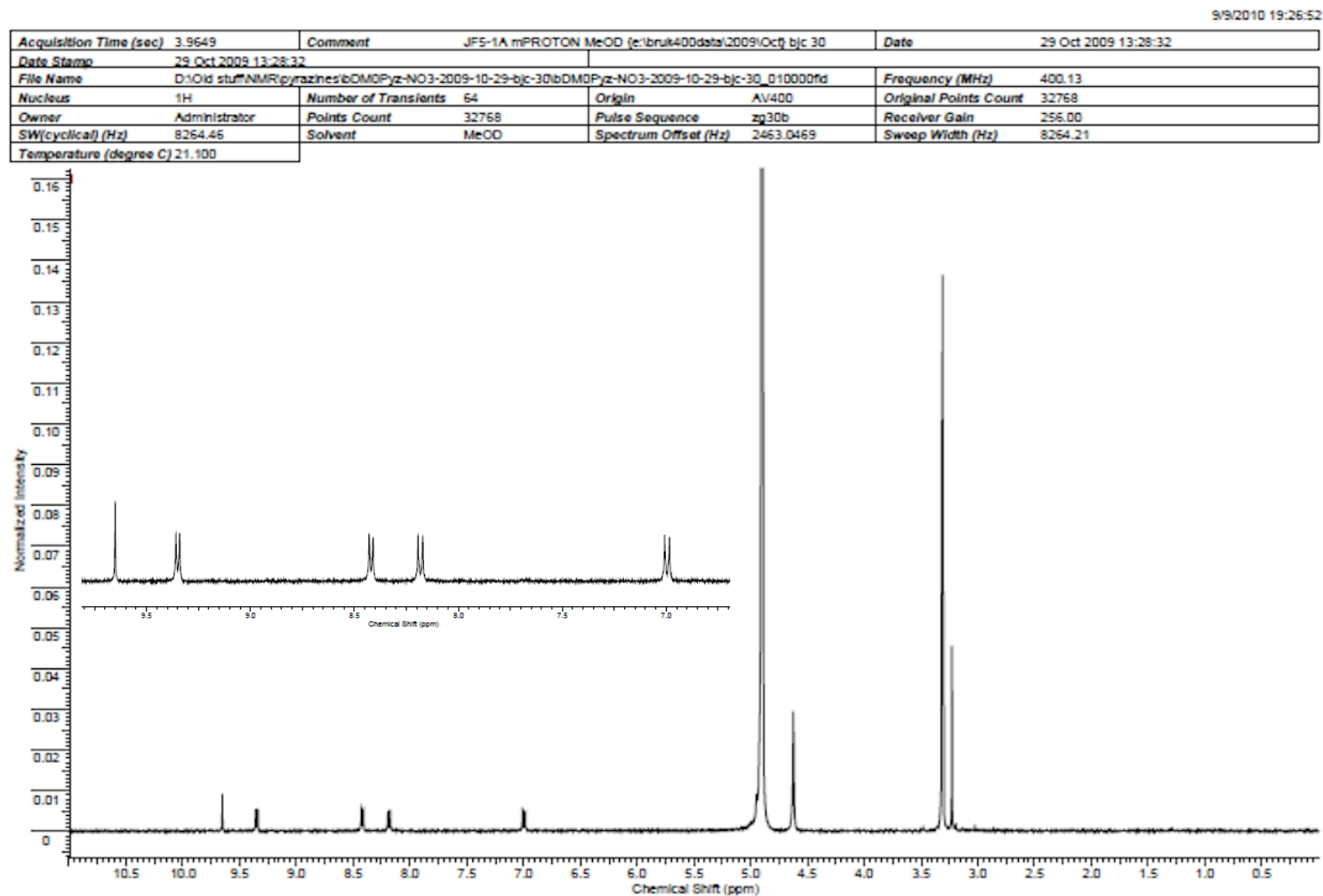
**Figure S6.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{2}][\text{PF}_6]_2$  (400 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).



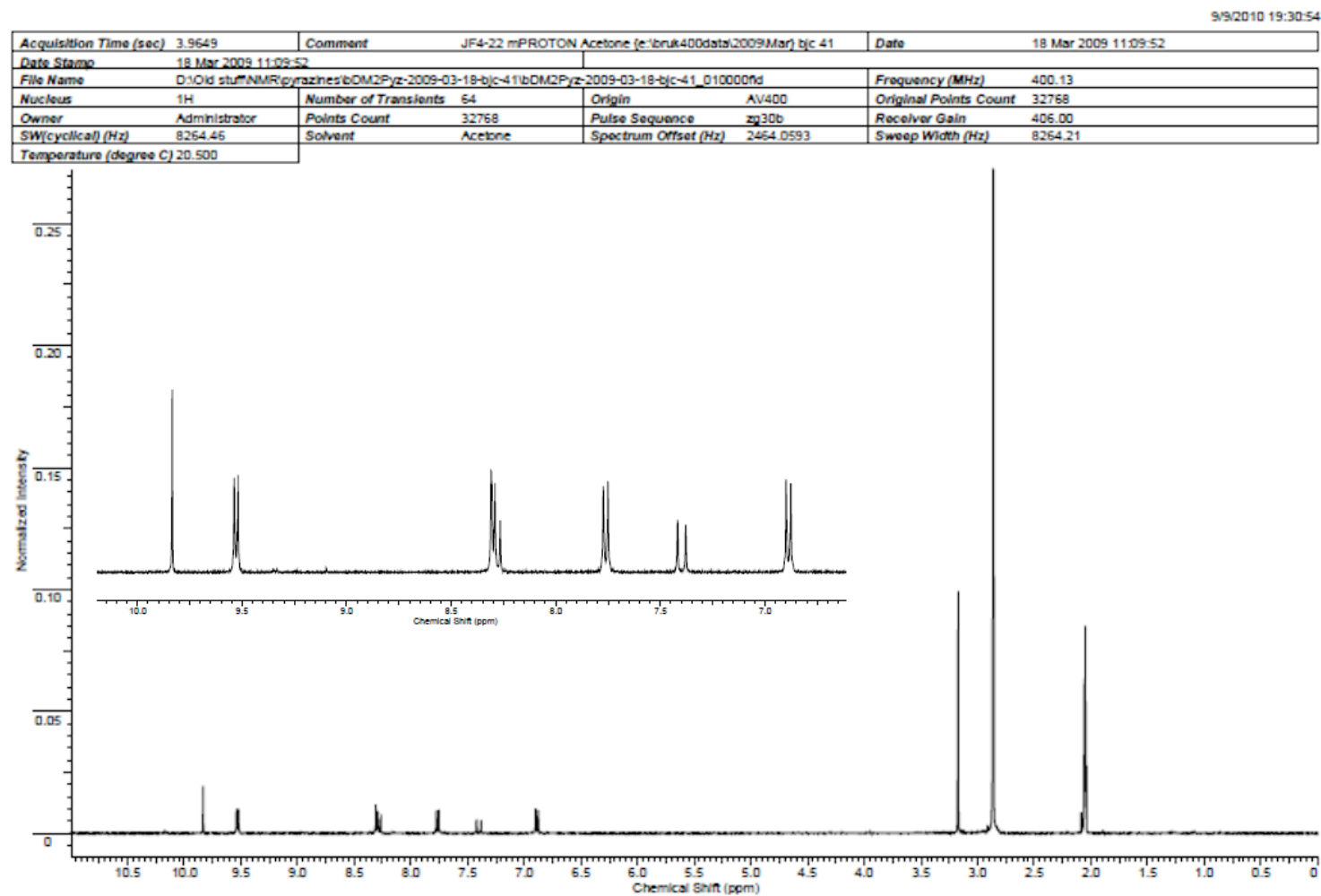
**Figure S7.**  $^{13}\text{C}$  NMR spectrum of **[2]** $[\text{PF}_6]_2$  (100 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K).



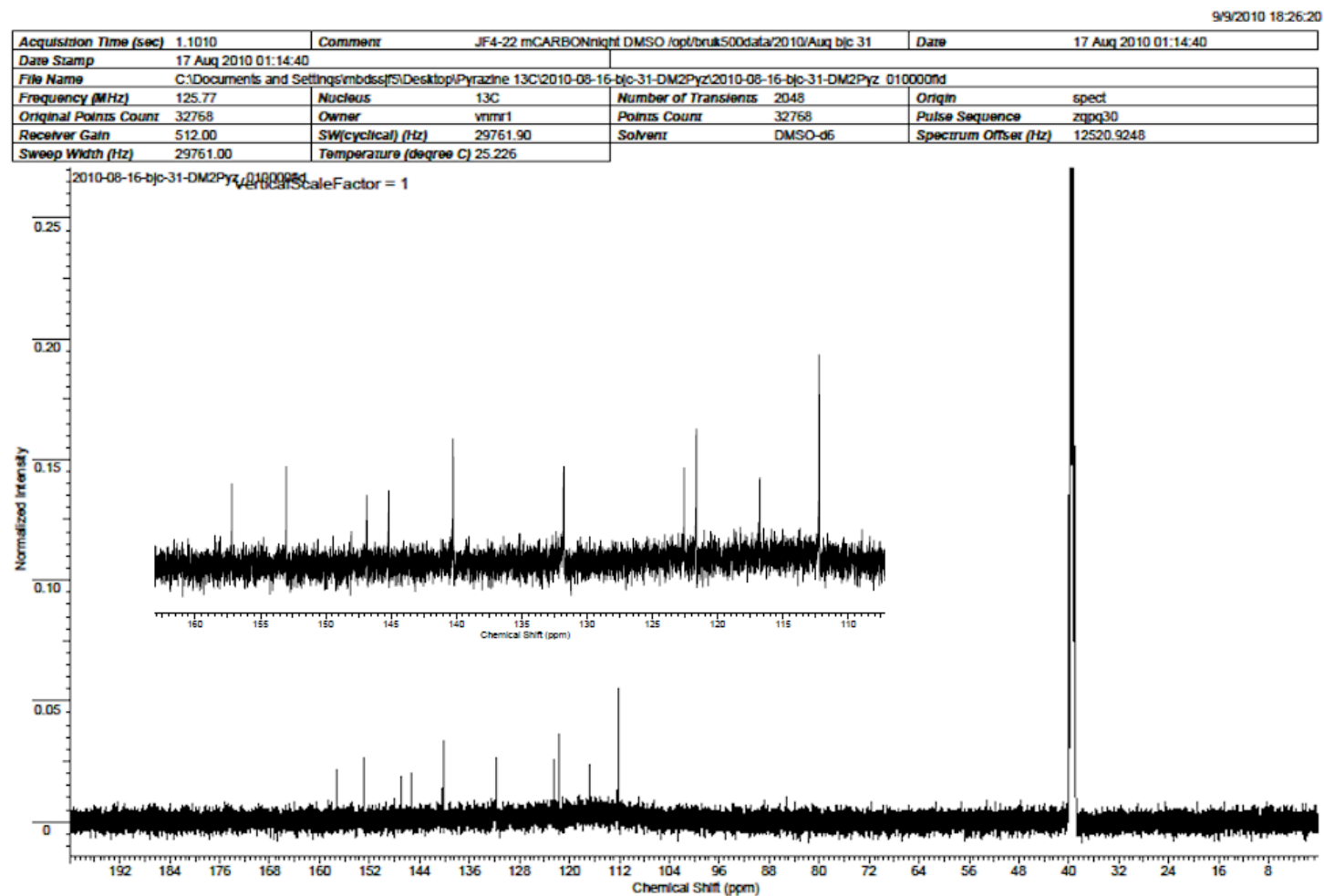


**Figure S8.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{2}][\text{NO}_3]_2$  (400 MHz in  $\text{CD}_3\text{OD}$  at 293 K).

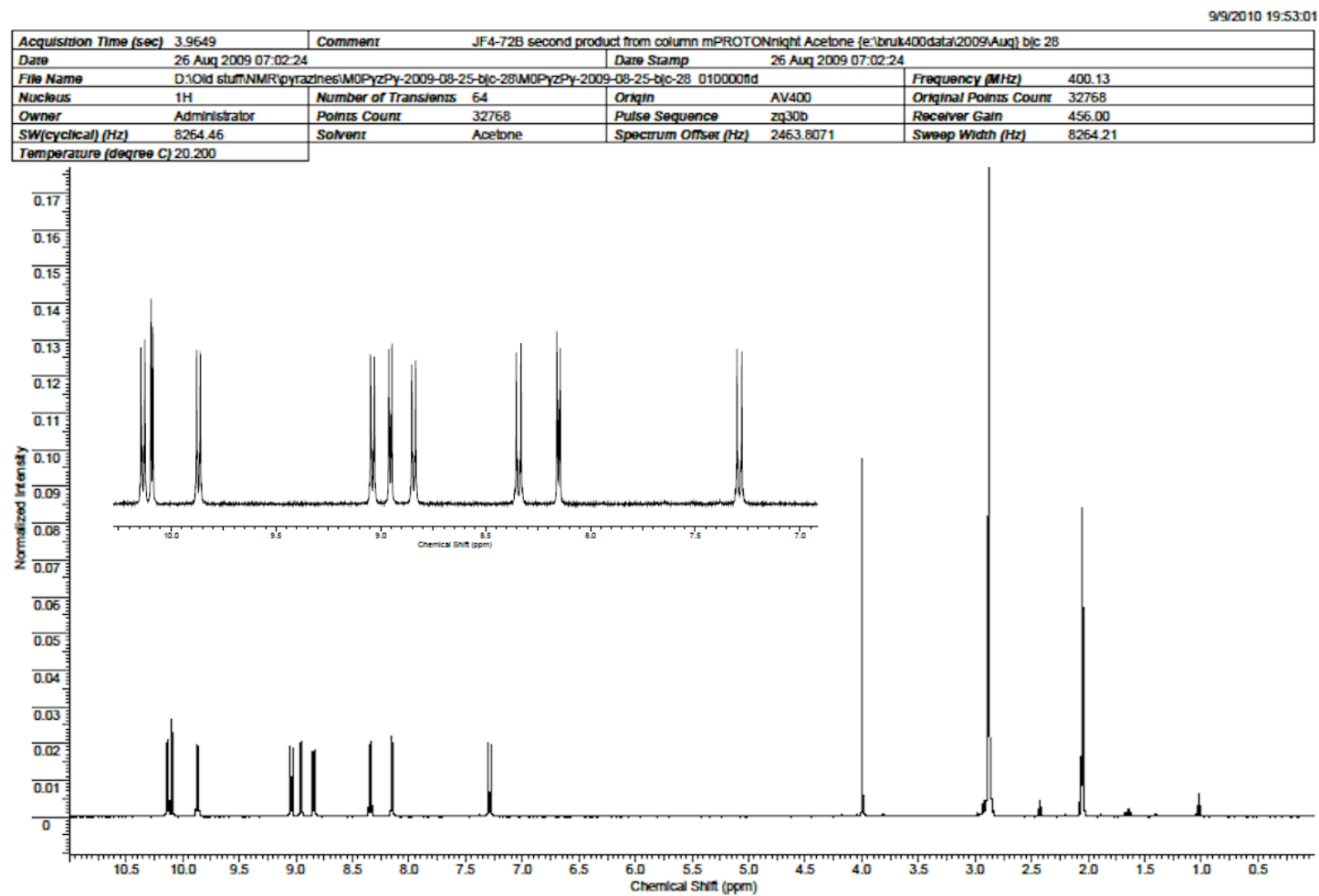
**Figure S9.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{3}][\text{PF}_6]_2$  (400 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).



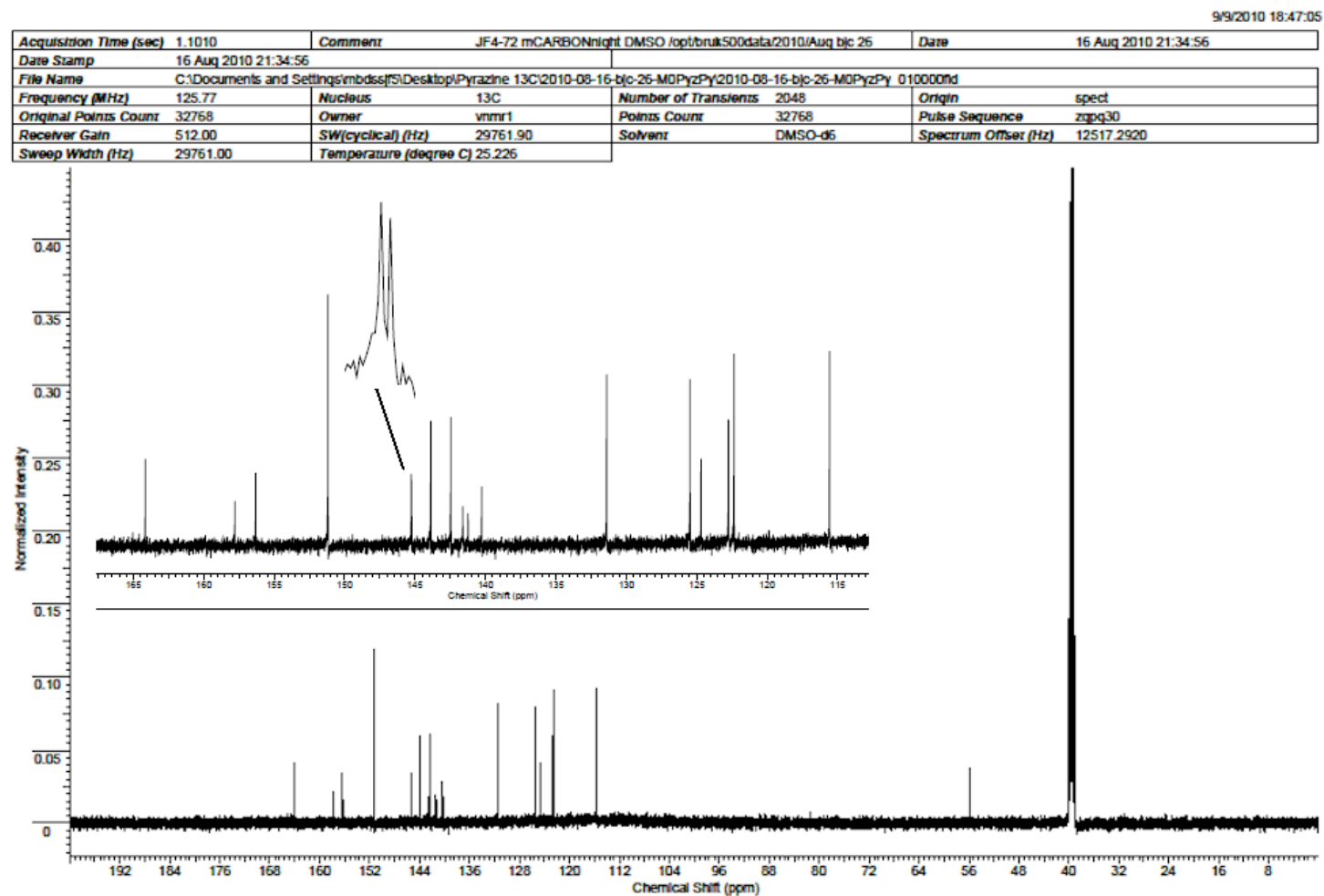
**Figure S10.**  $^{13}\text{C}$  NMR spectrum of  $[\mathbf{3}][\text{PF}_6]_2$  (125 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K).



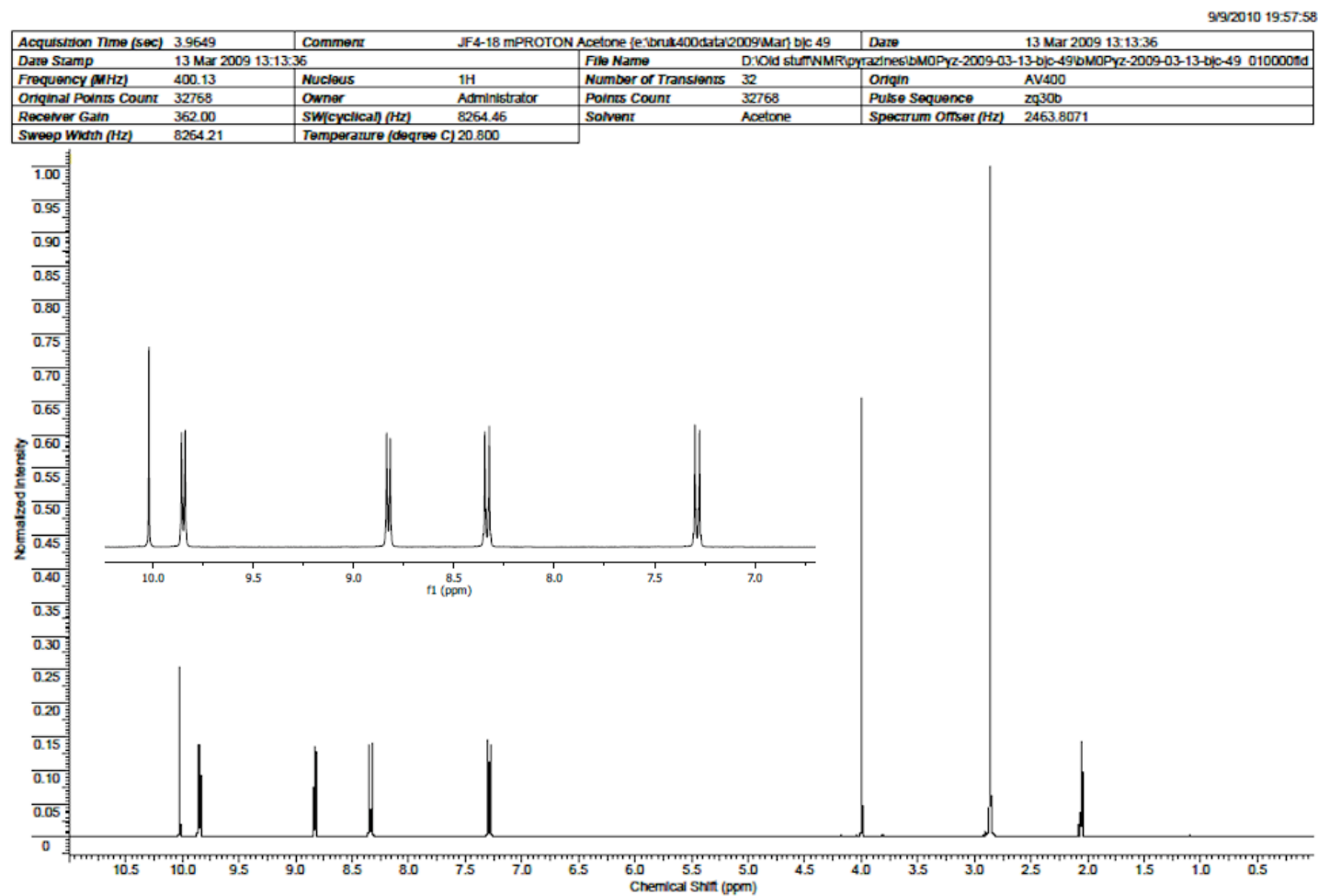
**Figure S11.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{4}][\text{PF}_6]_2$  (400 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).



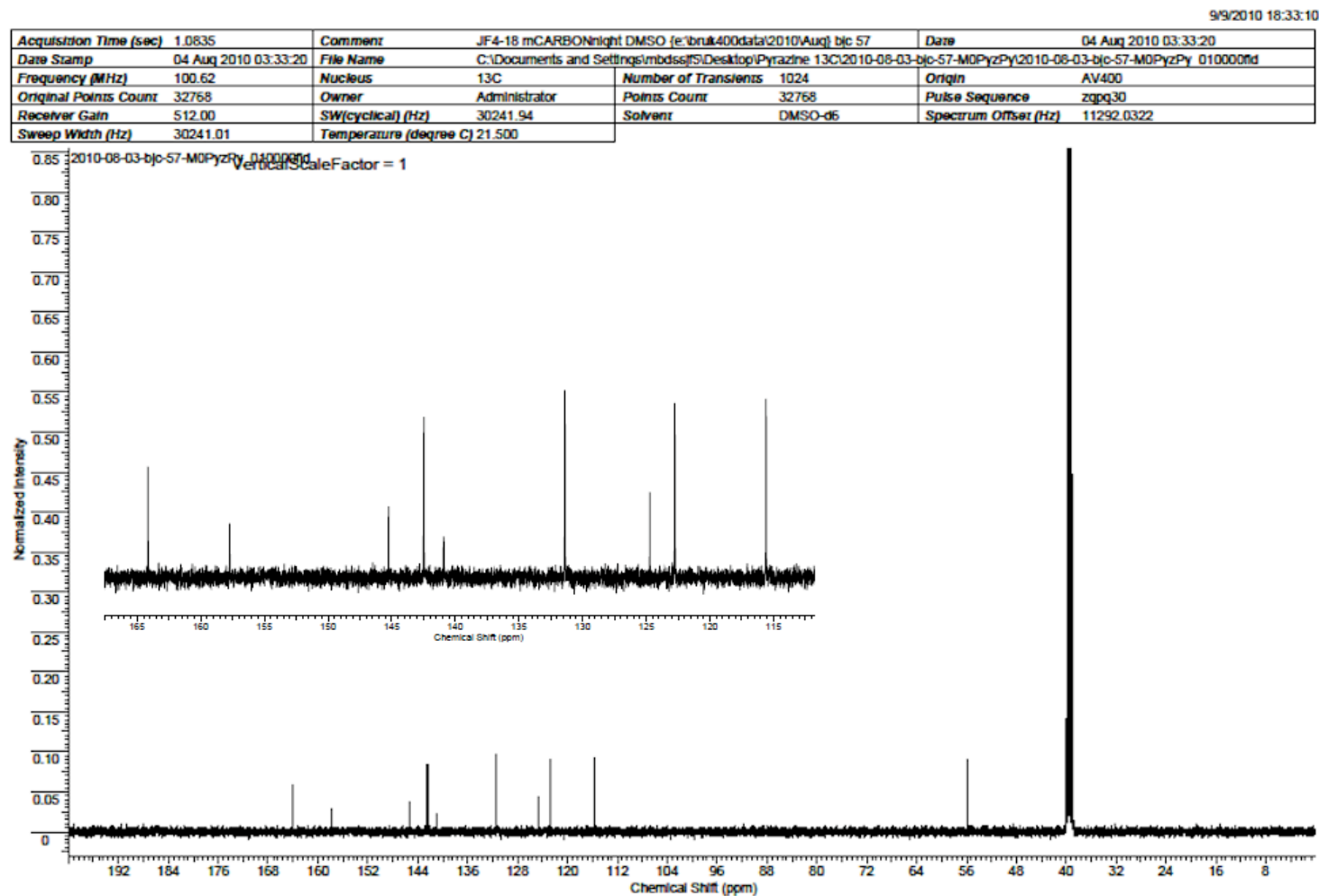
**Figure S12.**  $^{13}\text{C}$  NMR spectrum of  $[\mathbf{4}][\text{PF}_6]_2$  (125 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K).  
(extra expansion shows the closely spaced environments at 145.27/145.25 ppm)

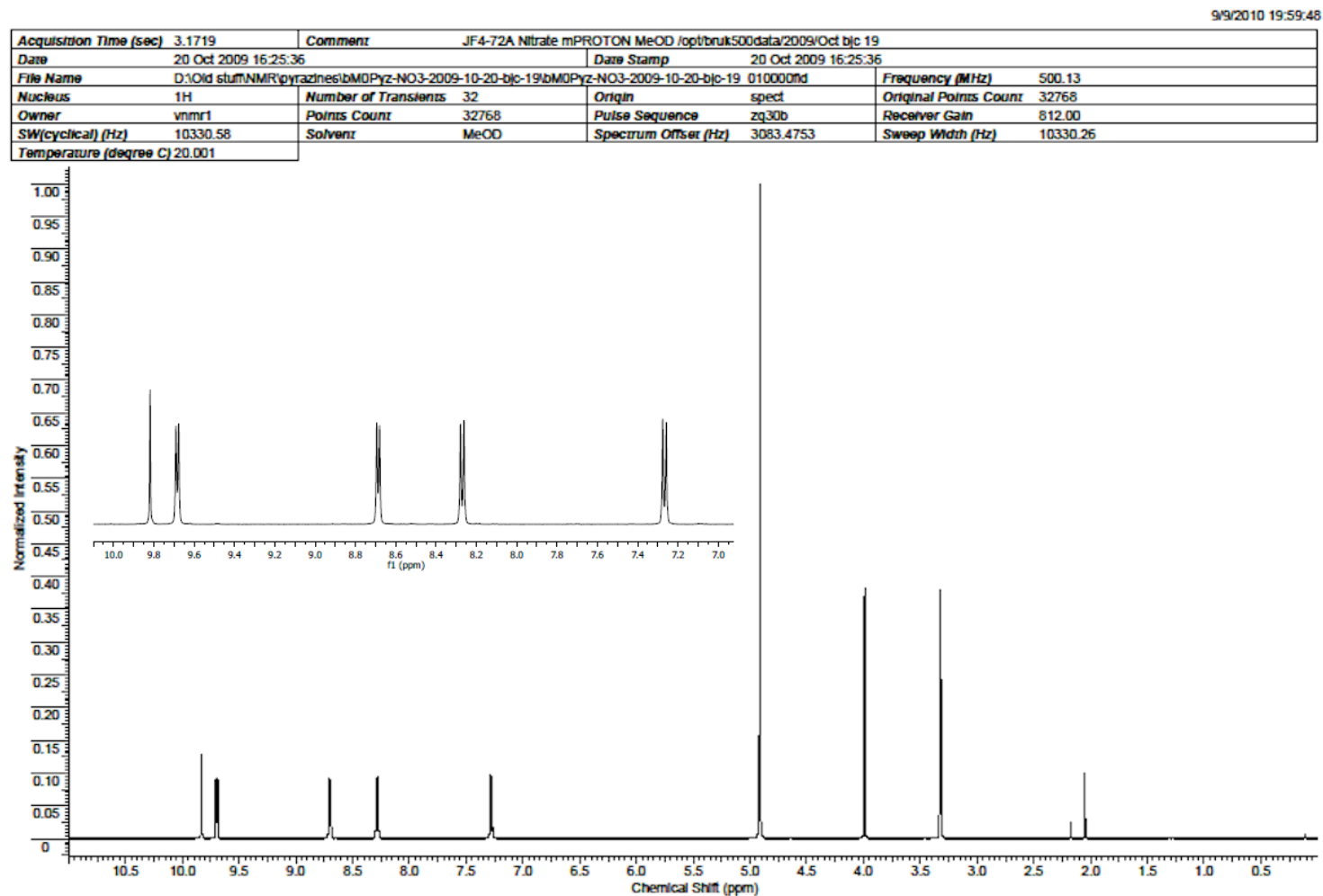


**Figure S13.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{5}][\text{PF}_6]_2$  (400 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).

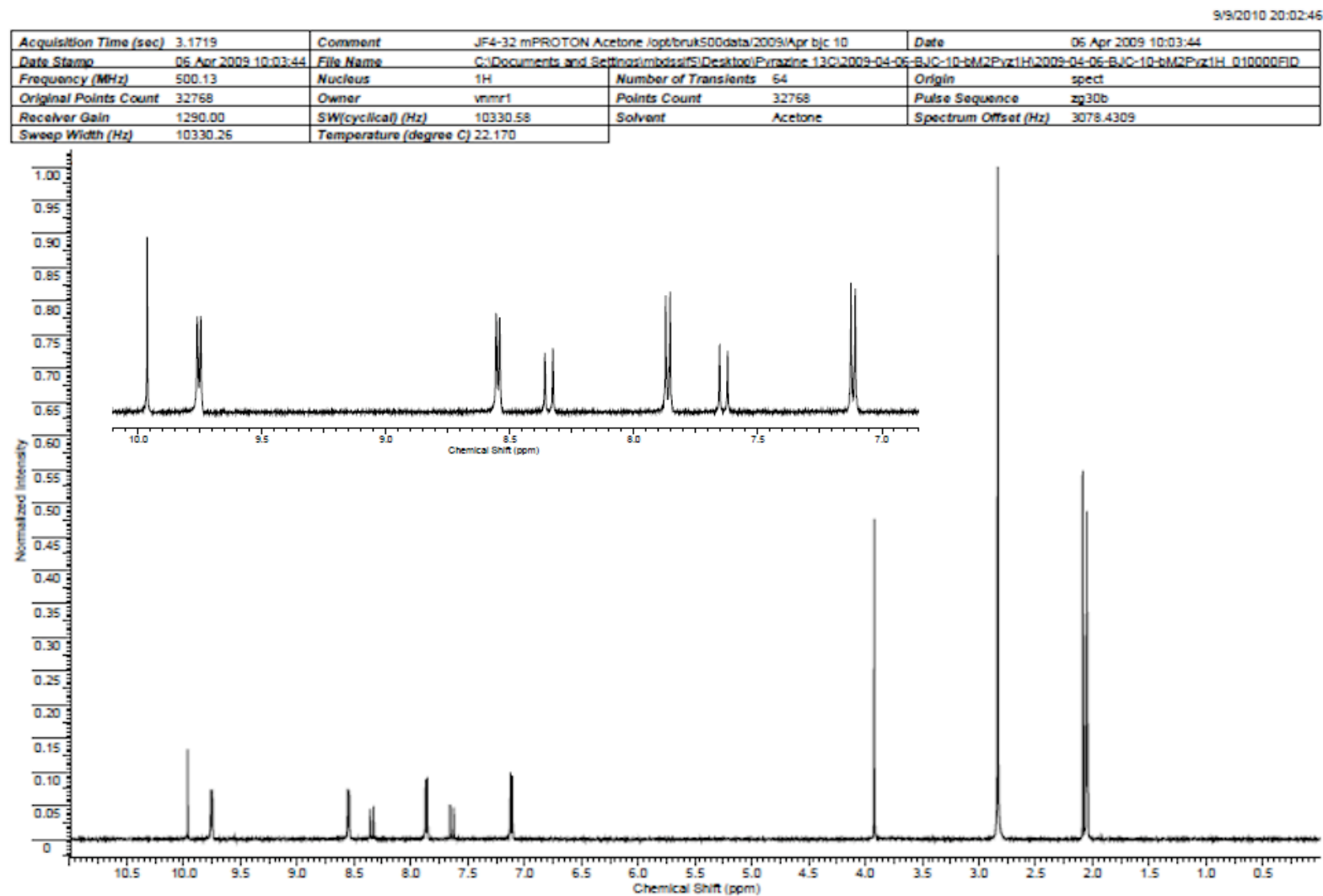


**Figure S14.**  $^{13}\text{C}$  NMR spectrum of  $[\mathbf{5}][\text{PF}_6]_2$  (100 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K).



**Figure S15.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{5}][\text{NO}_3]_2$  (500 MHz in  $\text{CD}_3\text{OD}$  at 293 K).



**Figure S16.**  $^1\text{H}$  NMR spectrum of  $[\mathbf{6}][\text{PF}_6]_2$  (500 MHz in  $(\text{CD}_3)_2\text{CO}$  at 293 K).

**Figure S17.**  $^{13}\text{C}$  NMR spectrum of  $[\mathbf{6}][\text{PF}_6]_2$  (75 MHz in  $(\text{CD}_3)_2\text{SO}$  at 293 K).

